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CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 9 CHEMISTRY WORKSHEET # 6 MIXED MOLE PROBLEMS (GRAMS, MOLECULES, AND LITERS) You now know three things a mole can be: a molar mass, 6.02×10^{23} molecules and, for a gas, 22.4 liters at STP.

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In each case we simply calculate the sum of the "weights" of the atoms in the formula to determine the weight of a mole. These weights can be found on the periodic table. EXAMPLE: Calculate the molar mass of a mole of iodine, I. 2. Round to 2 decimal places. $2 \text{ I} = 2 \times (126.90) = 253.80 \text{ g I. 2/mol}$.

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Question: PAGES CHEMISTRY B-MOLES PACKET NAME: HR: CHEMISTRY WORKSHEET #5 MOLE PROBLEMS-MOLAR VOLUME OF A GAS We Have

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Learned That A Mole Is A Mass Of Material And Number Of Particles. A Mole Can Also Be A Measure Of Volume When We Are Talking About Gases. You May Remember From Previous Science Classes That All Gases Behave Basically The Same As Far As The Physical ...

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CHEMISTRY B- MOLES PACKET NAME: _____ HR: _____ PAGE 6CHEMISTRY WORKSHEET # 4 MOLAR MASSES AND AVOGADRO'S NUMBERNow that you know two definitions of a mole (a molar mass and an Avogadro's number of particles) you can combine these two definitions into one problem.EXAMPLE: How many molecules are there in 90.1 grams of water? $2 \text{ H} = 2 \times (1.01) = 2.02$ $\text{O} = 1 \times (16.00) = 16.00$ $\text{H}_2\text{O} = 18.02 \text{ g? molecules}$ $\text{H}_2\text{O} = 90.1 \text{ g H}_2\text{O} \times 1 \text{ mole H}_2\text{O} \times 6.02 \times 10^{23} \text{ molecules H}_2\text{O} = 3.01 \times 10^{24} \text{ molecules H}_2 \dots$

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Packet Chemistry Answers

1 mole = 6.022×10^{23} particles or 602 200 000 000 000 000 000 000 particles (More than you could count in a lifetime!) This number is called

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Avogadro's number, named after Amedeo Avogadro.

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While a dozen is only 12 particles a mole is a much larger number— 6.02×10^{23} particles. Elements generally exist as the particles we call atoms. A mole of carbon contains 6.02×10^{23} atoms of carbon. A mole of helium contains 6.02×10^{23} atoms of helium. A mole of sodium contains 6.02×10^{23} atoms of sodium.

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Simply multiply this number of grams per mole times the number of moles you have in order to get the final answer: grams in 0.2 moles of $\text{CO}_2 = 0.2 \text{ moles} \times 44.01 \text{ grams/mole} = 8.80 \text{ grams}$ It's good practice to make certain units cancel out to give you the one you need.

What Is a Mole in Chemistry? - ThoughtCo

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