

## Concept Review How Substances Dissolve Answer Key

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### Concept Review How Substances Dissolve

How Substances Dissolve continued What Can Cause a Solute to Dissolve More Quickly? A solute dissolves because its particles interact with the particles of a solvent. Anything that allows more solvent to touch more solute will cause a solute to dissolve more quickly. INCREASE IN SURFACE AREA Small pieces of a substance dissolve faster than large pieces.

### 8 Solutions SECTION 2 How Substances Dissolve

The following substances all dissolve to some extent in water. Classify each as an electrolyte or a nonelectrolyte. acetone (CH<sub>3</sub>COCH<sub>3</sub>) iron(III) nitrate [Fe(NO<sub>3</sub>)<sub>3</sub>] elemental bromine (Br<sub>2</sub>) sodium hydroxide (NaOH) Answer. a. nonelectrolyte. b. electrolyte. c. nonelectrolyte. d. electrolyte

### 9.3: The Dissolution Process - Chemistry LibreTexts

Concept Review Section: How Substances Dissolve 1. Explain how you can speed up the dissolving process when preparing juice from frozen concentrate. \_\_\_\_\_ 2. Explain why water is sometimes referred to as the universal solvent. \_\_\_\_ ...

### Skills Worksheet Concept Review

The solubility in a substance is the maximum mass of a solute that can dissolve in 100g of solvent at a certain temperature and standard atmospheric atmosphere pressure. Some substances, such as acetic acid, methanol, ethanol, glycerol, and ethhylene glycol are completely souble in water.

### How Substances Dissolve Flashcards | Quizlet

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### Concept Review Section How Substances Dissolve Answers

The best way to tell if something will dissolve is to look at the polarities of the solvent and the solute. If the polarities of the solvent and solute match (both are polar or both are nonpolar), then the solute will probably dissolve.

### Chemistry: How and Why Do Things Dissolve?

Answers # concept review section how substances dissolve answers concept review section how substances dissolve 1 explain how you can speed up the dissolving process when preparing juice from frozen concentrate 2 explain why water is sometimes referred to as the universal solvent concept

### Concept Review Section How Substances Dissolve Answers

-to dissolve an ionic substance, water molecules must exert a force on the ions that is more attractive than the force between the ions in the crystal Hydrogen Bonding -hydrogen bonding is the intermolecular force occurring when a hydrogen atom that is bonded to a highly electronegative atom of one molecule is attracted to two unshared electrons of another molecule

### Section 2- How Substances Dissolve Flashcards | Quizlet

Water is considered a polar solvent. Which substances should dissolve in water? methanol (CH<sub>3</sub>OH) sodium sulfate (Na<sub>2</sub>SO<sub>4</sub>) octane (C<sub>8</sub>H<sub>18</sub>) Solution. Because water is polar, substances that are polar or ionic will dissolve in it. Because of the OH group in methanol, we expect its molecules to be polar. Thus, we expect it to be soluble in water.

### 9.1: Solutions - Chemistry LibreTexts

Each substance can be classified as an ionic solute or a nonionic solute. Ionic solutes are electrolytes, and nonionic solutes are nonelectrolytes. Potassium chloride is an ionic compound; therefore, when it dissolves, its ions separate, making it an electrolyte. Fructose is a sugar similar to glucose.

### 9.3 The Dissolution Process | The Basics of General ...

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### (PDF) Physical Science Concept Review Worksheets with ...

Review how students can change the rate at which substances dissolve with the true/false questions and skill challenge in this worksheet. Skip to main content ... Concepts and Challenges of Physical Science. From Pearson's Concepts and Challenges Physical Science.