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where: $\text{pH} = -\log_{10} a_{\text{H}^+}$ and. $\text{pOH} = -\log_{10} a_{\text{OH}^-}$. Equation 2 states that as pH increases pOH decreases correspondingly and vice versa because pK_w is constant for a given temperature. At 25°C, pH 7.0 is neutral, the activities of the hydrogen and hydroxyl ions are equal, and each

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corresponds to an approximate activity of 10^{-7} moles/L. The neutral point is temperature-dependent and is ...

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sample on which it is performed. It is beyond the scope of this publication to specify detailed procedures for the collection of all samples because of varied purposes and analytical procedures.

1060 COLLECTION AND PRESERVATION OF ... - Standard

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The analytical methods included in this part make use of classical wet chemical techniques and their automated variations and such modern instrumental techniques as ion chromatography. Methods that measure

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various forms of chlorine, nitrogen, and phosphorus are presented. The

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November 27, 2017 This notice is to
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are now available online. Below is a list of sections that underwent substantive changes between editions:

for - Yabesh

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Use 372 mg/L of the disodium salt of
ethylenediaminetetraacetic acid (EDTA).
Adjust EDTA solution to pH 6.5 before

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use. Add EDTA se parately to sample bottle before bottle sterilization (0.3 mL 15% solution in a 120-mL bottle) or combine it with the $\text{Na}_2\text{S}_2\text{O}_3$ solution before addition. 3.

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herein to the standard methods described elsewhere in this publication. Other sections of Part 1000 address laboratory equipment, laboratory safety, sampling procedures, and method

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TABLE OF CONTENTS Part 1000
INTRODUCTION 1010 INTRODUCTION A.
Scope and Application of Methods B.
Statistics C. Glossary D.
Dilution/Concentration Options 1020
QUALITY ASSURANCE A. Introduction B.
Quality Control C. Quality Assessment
1030 DATA QUALITY

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