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Section 9.2 1. For each of the following
pairs of elements, underline the one that

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would replace the other element in a compound. a. calcium, tin e. iron, copper
b. bromine, fluorine f. iodine, chlorine c.
aluminum, potassium g. silver, lead d.
zinc, sodium 2.

TEACHING TRANSPARENCY WORKSHEET The Activity Series

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by Teachers for Teachers and then utilized by teachers who are writing their own transparency sheets. These questions address topics like: Who does what?

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and 41

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Teaching Transparency 32 Chapter 9 Answer Key ...

Use with Chapter 3, Section 3.4 S (-QS 2.
What is the difference between a
mixture and a pure substance? Q 3.
What is the difference between a
compound and an element? 4. Can a
compound be a heterogeneous mixture?

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Explain your answer by referring to the
... Teaching Transparency Worksheets
Chemistry: Matter and Change Chapter 3

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CHAPTER Section 9.2 continued Date
Class STUDY GUIDE In your textbook,

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read about the activity series for metals and halogens. II on Examine each of the following pairs of potential reactants. Use Figure 9.13 in your textbook to help you decide whether or not a reaction would occur. If a reaction occurs, write the balanced equation.

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Worksheets: These transparencies relate to major concepts that will benefit from an extra visual learning aid. Most of the transparencies contain art or photos that extend the concepts put forth in the textbook. Others contain art or photos directly from the Student Edition. There are 92 Teaching Transparencies,

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UNIT 1 RESOURCES Earth Science

MATH SKILLS TRANSPARENCY

WORKSHEET Section 17 Use with

Chapter 11, 1.2 MATH SKILLS

TRANSPARENCY WORKSHEET 1. Explain

what a crystalline solid is. 18 Use with

Chapter 12, Section 2.3 1. The reaction

of iron(III) oxide (Fe_2O_3) and hydrogen

(H_2 is represented by the following

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unbalanced chemical equation. Fe_2O_3
(s) + H_2 (g) \rightarrow Fe (s) + H_2O (l)

MATH SKILLS TRANSPARENCY MASTER 16

Solutions Manual Chemistry: Matter and
Change • Chapter 9 143 CHAPTER 9
SOLUTIONS MANUAL Problem-Solving
Lab page 294 Halogen Atomic Radius

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(ppm) Ionic Energy (kJ/mol)

Electronegativity Fluorine 72 1681 3.98

Chlorine 100 1251 3.16 Bromine 114

1140 2.96 Iodine 133 1008 2.66 Astatine

140 920 2.2 Properties of Halogens 1.

Chemical Reactions

9. Is the outer electron configuration of

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the O ion after the reaction a very stable one? How can you tell? CD 10. What is the electron configuration of the Ca ion?
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